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GaBr_3

(a) The O.N. (oxidation number) of P decreases from 0 in P_4 to -3 in PH_3 and increases from 0 in P_4 to +2 in HPO_2 . Hence, P_4 acts both as an oxidizing agent and a reducing agent in this reaction.. Ion-electron method: The oxidation half equation is: $\text{P}_4(\text{s}) \rightarrow \text{H}_2\text{PO}_2^-(\text{aq})$. The P atom is balanced as:

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